Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Formula Writing Sheet #1

**Using ion reference sheet and notes complete the chart by writing the correct formulas for the following compounds.**

|  |  |  |  |
| --- | --- | --- | --- |
| Name of Compound | Positive ion | Negative ion | Formula |
| 1. Sodium Iodide | Na+ | I- | NaI |
| 1. Silver sulfide | Ag+ | S2- | Ag2S |
| 1. Barium sulfate |  |  |  |
| 1. Lithium sulfide |  |  |  |
| 1. Sodium hydroxide |  |  |  |
| 1. Ammonium chlorate |  | ClO3- |  |
| 1. Zinc sulfate | Zn2+ |  |  |
| 1. Iron(III) phosphate | Fe3+ |  |  |
| 1. Nickel (II) hydroxide | Ni2+ |  |  |
| 1. Chromium (III) oxide | Cr3+ |  |  |
| 1. Iron (III) sulfate |  |  |  |
| 1. Copper (II) nitrate |  |  |  |
| 1. copper (II) carbonate |  |  |  |
| 1. magnesium phosphide |  |  |  |
| 1. aluminum nitrate |  |  |  |
| 1. sodium phosphate |  |  |  |
| 1. aluminum sulfate |  |  |  |
| 1. aluminum sufide |  |  |  |
| 1. iron (III) sulfite |  |  |  |
| 1. ammonium carbonate |  |  |  |

# Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Formula Writing/Counting Atoms 2

Complete the chart. Remember these rules:

1. The number of atoms of each element (or polyatomic ions) is written in the space below the line and to the right of the symbol as a subscript.
2. When the number of atoms (or polyatomic ions) is one, the one is “understood” and you do not write anything.
3. The positive atom (or polyatomic ion) is written first in the formula.
4. Use (parentheses) only when necessary.

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **Compound made of** | **Positive ion** | **Negative ion** | Formula | Compound Name | Number of atoms in compound |
| 1. calcium and   nitrate |  |  |  |  |  |
| 1. tin (IV) and   chloride |  |  |  |  |  |
| 1. copper (II) and   carbonate |  |  |  |  |  |
| 4. barium and  bromide |  |  |  |  |  |
| 1. tin (II) and   sulfite |  |  |  |  |  |
| 1. Ammonium and   Nitrate |  |  |  |  |  |
| 1. Lithium and   phosphorus |  |  |  |  |  |
| 1. Sodium and   Bicarbonate |  |  |  |  |  |
| 1. Lead (II) and   Phosphate |  |  |  |  |  |
| 1. magnesium and   hydroxide |  |  |  |  |  |
| 1. silver and   sulfide |  |  |  |  |  |
| 1. barium and   acetate |  |  |  |  |  |
| 1. fluorine and   manganese (II) |  |  |  |  |  |
| 1. Chromium (III)   and nitrate |  |  |  |  |  |
| 1. sulfate and   Iron (III) |  |  |  |  |  |

# Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Formula Writing/Counting Atoms 3

Complete the chart. Remember these rules:

1. The number of atoms of each element (or polyatomic ions) is written in the space below the line and to the right of the symbol as a subscript.
2. When the number of atoms (or polyatomic ions) is one, the one is “understood” and you do not write anything.
3. The positive atom (or polyatomic ion) is written first in the formula.
4. Us (parentheses) only when necessary.

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **Compound made of** | **Positive ion** | **Negative ion** | Formula | Compound Name | Number of atoms in compound |
| 1. calcium and   chlorate |  |  |  |  |  |
| 1. nickel (II) and   sulfate |  |  |  |  |  |
| 1. copper (I) and   carbonate |  |  |  |  |  |
| 4. magnesium and  chlorine |  |  |  |  |  |
| 1. tin (II) and   sulfate |  |  |  |  |  |
| 1. Ammonium and   phosphate |  |  |  |  |  |
| 1. aluminum and   nitrate |  |  |  |  |  |
| 1. calcium and   sulfite |  |  |  |  |  |
| 1. iron (III) and   carbonate |  |  |  |  |  |
| 1. calcium and   hydroxide |  |  |  |  |  |
| 1. lithium and   sulfate |  |  |  |  |  |
| 1. strontium and   carbonate |  |  |  |  |  |
| 1. fluorine and   ammonium |  |  |  |  |  |
| 1. Chromium (III)   and oxide |  |  |  |  |  |
| 1. phosphate and   Iron (II) |  |  |  |  |  |